#### Chemistry - Topic 15 - Transition Metals

TRANSITION METAL & elements that form one or more stable ions with immuniciply filled O-craits MOAND & a species with one or more incre pairs greatrons that surrounds a transition metal ion using coordinate bonding to form a complex ion

#### THE 3d SERIES OF THE PERIODIC TABLE

c) you need to know how to work out the electronic consiguration for any atom or ion.

SCANDIUM and ZING only form one ion such and their ions have no incomplete Dark cals.

Co SHOP THE P. North Total was Bell " may had to March Staff 95" 34" Section 18 of the

#### GENERAL IDEAS ABOUT TRANSITION METALS

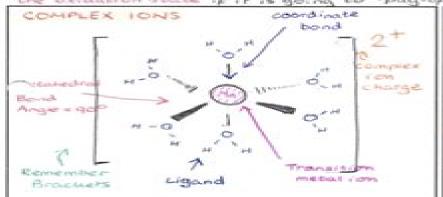
- \* They are hard solids with high melting and balling temperatures
- \* Can have various cicidation numbers so they make good catalusts
- A They form coloured ions when in solution as they absorb some lights.

REMEMBER: Copper and Chromium disobey the author principle

#### Mn= = 450 3d6 February and F Fe Ly = he Bd b , it is simply more aboute por the electrons to pill like this . N:2+ = 450 3d2 REMEMBER: Electrons from the 4s orbitos are lost before those in 3d!

#### WHY DO TRANSITION METALS HAVE VARIABLE OXIDATION STATES?

The compound formed is the one where the most energy is released as the more energy that is released, the more sponse the compound becomes. You need lots of energy to conside the metal but then more energy is released in hydration or lattice permation. As most of the electrons being removed in bransition metals, are in the O-orbital, it is not too costly to increase the avidation state if it is going to spay-off during lattice formation -



COMMON	MORE INFORMATION
Water	There are how lone pairs of
HzO	Whiter often forms octubed rows
Ammonia NH <sub>3</sub>	One lone pour of exectrons on the nutrogen above. Ammonia is small so forms octahedral structures.
Hydronde Ion OH	Forms hydroxo complexes, charge on the ion was reduce the charge of the complex ion.
Chloride /cm	form betrahedral completes -

OTHER LONS

V " = 45 3d =

#### HORE ON VARIABLE OXIDATION STATES

Transition metals can donate and recieve electrons and are able to act as acidising agents and as been reducing agence. The sub-shells of the D-orbital can easily gain or lose electrons meaning the energy disperences between the oxidation states is small.

- +2 state stability with respect to the 3+ state increases across a period.
- compounds with high acidation states tend to be acidising agents eig. MnOst (Mn is 7")
- compounds with low acadation states hend to be reducing agents e.g. V2+, Fe2+ (2+and 2

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