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 pK_b of the weak acid or base being titrated a solution ph as a function of the volume of 1.00 M NaOH added to 10.00 ml of
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web where the salt is the conjugate base of the acid the four parts of the titration curve are described below and you should
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web solution a titrant volume 0 ml the solution ph is due to the acid ionization of HCl because this is a strong acid the
ionization is complete and the hydronium ion molarity is 0.100 M the ph of the solution is then $ph = -\log 0.100 = 1.000$ $ph = -\log 0.100 = 1.000$ b titrant volume 12.50 ml

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reacting it with a standard solution one type of titration uses a neutralization reaction in which an acid and a base react to
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web at the equivalence point in an acid base titration moles of base = moles of acid and the solution only contains salt and
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web expert answer 100% 2 ratings the example is based upon henderson hasselbalch equation the equation gives relation

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Volume of NaOH added (mL)	pH
4.6	4.6
4.6	5.9
5.9	5.9

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web background plot of the ph of a solution against the volume of titrant added is called a titration curve the ph can be measured directly with a ph meter while titrant is added from a buret for the acids used in this experiment the titrant will always be a 0.1 M solution of the strong base NaOH

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