Table 6.3 Standard Enthalpies of Formation of Some Inorganic Substances at 25°C

Substance	ΔH ^o (kJ/mol)	Substance	ΔH ² (kJ/mol)
Ag(s)	0	H ₂ O ₂ (I)	-187.6
AgCl(s)	-127.04	Hg(/)	0
Al(s)	0	I ₂ (s)	0
Al ₂ O ₃ (s)	-1669.8	HI(g)	25.94
Br ₂ (/)	0	Mg(s)	0
HBr(g)	-36.2	MgO(s)	-601.8
C(graphite)	0	MgCO ₃ (s)	-1112.9
C(diamond)	1.90	N₂(g)	0
CO(g)	-110.5	NH₃(g)	-46.3
$CO_2(g)$	-393.5	NO(g)	90.4
Ca(s)	0	NO ₂ (g)	33.85
CaO(s)	-635.6	N ₂ O ₄ (g)	9.66
CaCO ₃ (s)	-1206.9	N₂O(g)	81.56
Cl ₂ (g)	0	O(g)	249.4
HCI(g)	-92.3	O ₂ (g)	0
Cu(s)	0	O ₃ (g)	142.2
CuO(s)	-155.2	S(rhombic)	0
F ₂ (g)	0	S(monoclinic)	0.30
HF(g)	-268.61	SO₂(g)	-296.1
H(g)	218.2	SO₃(g)	-395.2
H ₂ (g)	0	H₂S(g)	-20.15
$H_2O(g)$	-241.8	ZnO(s)	-347.98
H ₂ O(/)	-285.8		

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LL Leslie



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