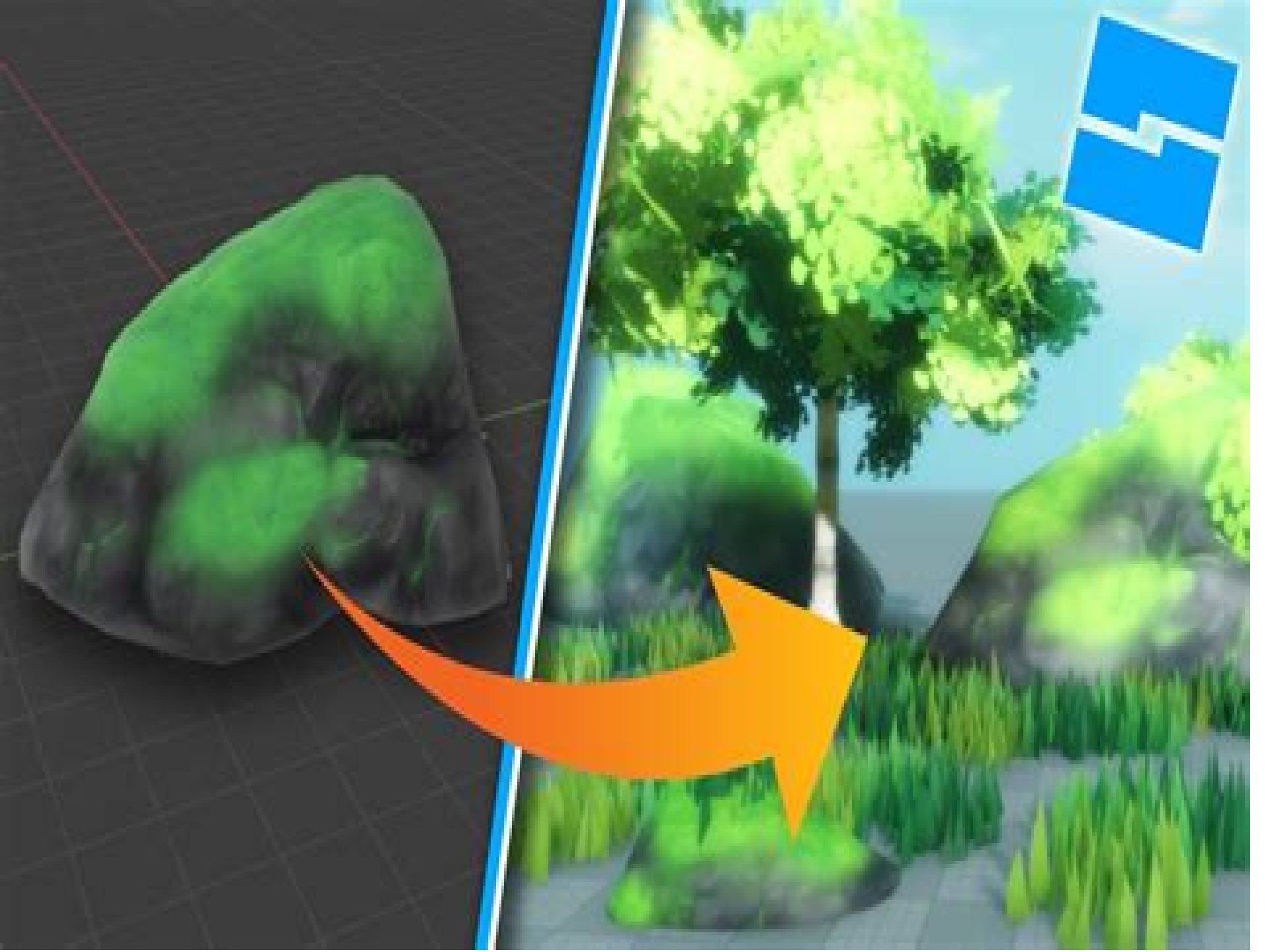




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web limiting reactant theoretical and percentage yields key $4\text{KO}_2 + 2\text{H}_2\text{O} \rightarrow 4\text{KOH} + 3\text{O}_2$ If a reaction vessel contains 0 mol KO_2 and 0 mol H_2O , what is the limiting reactant?

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 $\text{Cu} + 2\text{AgNO}_3 \rightarrow \text{Cu(NO}_3)_2 + 2\text{Ag}$

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 $3\text{CO} + 7\text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3\text{H}_2\text{O}$

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web start with 8.3 mole H_2S and multiply it with 2 mole H_2O over 2 mole H_2S from the equation $8.3 \text{ mole } \text{H}_2\text{S} \times \frac{2 \text{ mole } \text{H}_2\text{O}}{2 \text{ mole } \text{H}_2\text{S}}$

mole H_2S x 18g H_2O i mole H_2O 149 4g H_2O is your theoretical yield so percent yield $\frac{137}{149} \times 100 = 91.77\%$ problems 6 8 are similar to problem 5

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web a write a balanced equation for the reaction $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$ b determine the theoretical yield of Al_2O_3 160.0 g Al x $\frac{1 \text{ mol Al}_2\text{O}_3}{2 \text{ mol Al}} \times 101.96 \text{ g Al}_2\text{O}_3 = 81.2 \text{ g Al}_2\text{O}_3$ c determine the percent yield $\frac{66.0 \text{ g}}{81.2 \text{ g}} \times 100 = 81.3\%$

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