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For This Character?

Bruno Is Sorry

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web an oxidation reduction redox reaction is one in which one or more atoms change oxidation numbers oxidation occurs when an atom s oxidation state becomes more positive indicating that electrons have been lost reduction occurs when an atom s oxidation state becomes less positive indicating that electrons have been gained

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web describe redox reactions in terms of electron transfer recall that oxidation involves loss of electrons and reduction involves gain of electrons represent simple redox reactions using half equations identify oxidising and reducing agents for simple redox reactions understand the reactivity of metals and metal displacement reaction

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what do the following terms mean oxidation oxidation is a loss of electrons reduction reduction is a gain of electrons what is a helpful way to remember this

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web write the oxidation and reduction reactions for each redox reaction the first one is done for you $23 \text{ Fe} + \text{Co} \rightarrow \text{Co}^{2+} + \text{Fe}^{2+}$
 oxidation $\text{Co} \rightarrow \text{Co}^{2+} + 2\text{e}^-$ reduction $\text{Fe}^{2+} + 2\text{e}^- \rightarrow \text{Fe}$ $24 \text{ 3 Ag} + \text{Ni}^{2+} \rightarrow \text{Ni} + 3 \text{ Ag}^{+}$ oxidation reduction $25 \text{ Cu} + \text{Pb}^{2+} \rightarrow \text{Pb} + 2 \text{ Cu}^{+}$ oxidation reduction 26 O_2

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 $\text{a } \text{I}_2 + \text{O}_5 \rightarrow \text{I}_2\text{O}_5$
 $5\text{CO} + \text{I}_2 \rightarrow 5\text{CO}_2 + \text{I}_2$ $\text{O}_5 + 5\text{C} + \text{O}_2 \rightarrow 5\text{CO}_2$

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