

Research Article

Third-Order Approximate Solution of Chemical Reaction-Diffusion Brusselator System Using Optimal Homotopy Asymptotic Method

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The objective of this paper is to investigate the effectiveness and performance of optimal homotopy asymptotic method in solving a system of nonlinear partial differential equations. Since mathematical modeling of certain chemical reaction-diffusion experiments leads to Brusselator equations, it is worth demanding a new technique to solve such a system. We construct a new efficient recurrent relation to solve nonlinear Brusselator system of equations. It is observed that the method is easy to implement and quite valuable for handling nonlinear system of partial differential equations and yielding excellent results at minimum computational cost. Analytical solutions of Brusselator system are presented to demonstrate the viability and practical usefulness of the method. The results reveal that the method is explicit, effective, and easy to use.

1. Introduction

The Brusselator model, the nonlinear system of partial differential equations, arises in the modeling of certain chemical reaction-diffusion processes. This Brusselator reaction-diffusion model plays a substantial role in the study of cooperative processes of chemical kinetics. This system occurs in a large number of physical problems. It arises in the creation of ozone by atomic oxygen through a triple collision, in enzymatic reactions, and in plasma and laser physics in numerous coupling among models [1]. A pair of variables are involved in dealing with these chemical reactions; intermediates with input and output chemicals, whose concentrations are likely to be controlled during the reaction process, are substantial under quite genuine conditions and are discussed by Nicolis and Prigogine in [2, 3]. This model has been revealed as the trimolecular model [4].

The two-dimensional nonlinear reaction-diffusion Brusselator system is

$$\begin{aligned}\frac{\partial}{\partial t} u(x, y, t) &= u^2 v - (A + 1)u + \mu \left(\frac{\partial^2 u}{\partial x^2} + \frac{\partial^2 u}{\partial y^2} \right) + B, \\ \frac{\partial}{\partial t} v(x, y, t) &= -u^2 v + Au + \mu \left(\frac{\partial^2 v}{\partial x^2} + \frac{\partial^2 v}{\partial y^2} \right),\end{aligned}\quad (1)$$

subject to the initial condition:

$$\begin{aligned}u(x, y, 0) &= h(x, y), \\ v(x, y, 0) &= g(x, y),\end{aligned}\quad (2)$$

where $u(x, y, t)$, $v(x, y, t)$ are unknown functions representing the dimensionless concentrations of two reactants, x , y , and t denote spatial and temporal independent variables, respectively, A and B are constant concentrations of the two reactants, μ represent the diffusion coefficient, and h and k are known functions. It is evident that, for small values of diffusion coefficient μ , steady state solution of Brusselator system converges to the equilibrium point $(B, A/B)$ if $1 - A + B^2 > 0$. During the last few years, the researchers have been interest in the existence of solution of the Brusselator reaction model when $1 - A + B^2 \neq 0$ [5–7]. In this paper, we have made a successful attempt to find the solution of such types of Brusselator system.

Numerical methods need large size of computational works and generally the consequence of round-off error causes loss of precision in the results for system of nonlinear partial differential equations. Analytical methods mostly used for solving these equations are very restricted and can

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